

(19)



Europäisches Patentamt

European Patent Office

Office européen des brevets



(11)

EP 1 415 972 A1

(12)

EUROPEAN PATENT APPLICATION

published in accordance with Art. 158(3) EPC

(43) Date of publication:

06.05.2004 Bulletin 2004/19

(51) Int Cl.7: **C07C 43/23, C07C 41/26**

(21) Application number: **02746040.1**

(86) International application number:

PCT/JP2002/007147

(22) Date of filing: **15.07.2002**

(87) International publication number:

WO 2003/006412 (23.01.2003 Gazette 2003/04)

(84) Designated Contracting States:

**AT BE BG CH CY CZ DE DK EE ES FI FR GB GR
IE IT LI LU MC NL PT SE SK TR**

Designated Extension States:

AL LT LV MK RO SI

(72) Inventors:

- **UEDA, Takahiro**
Kobe-shi, Hyogo 655-0872 (JP)
- **KITAMURA, Shiro**
Akashi-shi, Hyogo 673-0882 (JP)
- **UEDA, Yasuyoshi**
Himeji-shi, Hyogo 671-1227 (JP)

(30) Priority: **13.07.2001 JP 2001214482**

17.04.2002 JP 2002114877

(74) Representative: **HOFFMANN EITLE**

**Patent- und Rechtsanwälte
Arabellastrasse 4
81925 München (DE)**

(71) Applicant: **KANEKA CORPORATION**
Osaka-shi, Osaka 530-8288 (JP)

(54) METHOD OF PRODUCING REDUCED COENZYME Q10

(57) The present invention relates to a method of conveniently and efficiently producing high-quality reduced coenzyme Q₁₀ which is useful as an ingredient in foods, functional nutritive foods, specific health foods, nutritional supplements, nutrients, animal drugs, drinks, feeds, cosmetics, medicines, remedies, preventive drugs, etc. This method is suitable for industrial production thereof.

A method of producing a reduced coenzyme Q₁₀ which comprises reducing an oxidized coenzyme

Q₁₀ in an aqueous medium with the use of hyposulfurous acid or a salt thereof,

said reduction being carried out in the coexistence of a salt and/or under deoxygenated atmosphere, and at pH of 7 or below. Thus, the formation of the oxidized coenzyme Q₁₀ as a by-product due to oxidation can be minimized, thereby giving reduced coenzyme Q10 having excellent qualities in a high yield.

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Description

TECHNICAL FIELD

5 [0001] The present invention relates to a method of producing a reduced coenzyme Q₁₀. Reduced coenzyme Q₁₀ shows a higher level of oral absorbability as compared with oxidized coenzyme Q₁₀ and is a compound useful as an ingredient in good foods, functional nutritive foods, specific health foods, nutritional supplements, nutrients, animal drugs, drinks, feeds, cosmetics, medicines, remedies, preventive drugs, etc.

10 BACKGROUND ART

[0002] It is known that reduced coenzyme Q₁₀ can be prepared by producing coenzyme Q₁₀ in the conventional manner, for example by synthesis, fermentation, or extraction from natural products, and concentrating a reduced coenzyme Q₁₀-containing eluate fraction resulting from chromatography (JP-A-10-109933). On that occasion, as described in the above-cited publication, the chromatographic concentration may be carried out after reduction of oxidized coenzyme Q₁₀ contained in the reduced coenzyme Q₁₀ with a reducing agent such as sodium borohydride or sodium dithionite (sodium hyposulfite), or reduced coenzyme Q₁₀ may be prepared by reacting the reducing agent mentioned above with an existing highly pure grade of coenzyme Q₁₀.

15 [0003] JP-A-57-70834 discloses an example in which reduced coenzyme Q₁₀ was synthesized by dissolving coenzyme Q₁₀ in hexane and adding an aqueous solution of sodium hydrosulfite (sodium hyposulfite) in an amount of twice the weight of coenzyme Q₁₀ to the solution, followed by stirring.

[0004] However, the present inventors preliminary investigated the above reduction method, and found that it is not so easy to obtain a high-quality reduced coenzyme Q₁₀ in a high yield.

20 [0005] The above problem leads to not only economical disadvantageous but also to problems in qualities such as the immixture of oxidized coenzyme Q₁₀, which is difficult to remove, into a product. Moreover, use of a large amount of a reducing agent enhances the load for removal and detoxification of the reducing agent and components derived therefrom.

[0006] Thus, the above disadvantages in the reduction reaction give rise to a necessity of another process for purification.

30 SUMMARY OF THE INVENTION

[0007] In view of the foregoing, the present invention has an object to provide a convenient and efficient synthesis method to obtain high-quality reduced coenzyme Q₁₀.

35 [0008] The present inventors made intensive investigations, and as a result, found that high-quality reduced coenzyme Q₁₀ can be obtained at a high yield in a convenient and efficient manner, by carrying out a reduction reaction under specific condition in a method of producing reduced coenzyme Q₁₀ comprising reducing oxidized coenzyme Q₁₀ with hyposulfurous acid or a salt thereof. Based on this finding, the present inventors have completed the present invention.

40 [0009] Accordingly, the present invention is a method of synthesizing a reduced coenzyme Q₁₀ which comprises reducing an oxidized coenzyme Q₁₀ in an aqueous medium with the use of hyposulfurous acid or a salt thereof,

said reduction being carried out in the coexistence of a salt and/or under deoxygenated atmosphere, and at pH of 7 or below.

45 DETAILED DESCRIPTION OF THE INVENTION

[0010] Hereafter, the present invention is described in detail.

50 [0011] In the present invention, hyposulfurous acid or a salt thereof is used as a reducing agent. The hyposulfurous acid or a salt thereof is not particularly restricted but a salt form of hyposulfurous acid is generally used. The salt of hyposulfurous acid is not particularly restricted but includes, as preferred species, alkali metal salts, alkaline earth metal salts, ammonium salts and the like. Alkali metal salts such as a lithium salt, a sodium salt, and a potassium salt are more preferred, and a sodium salt is most preferred.

55 [0012] The above reduction reaction is carried out in an aqueous medium. The amount of water to be used in the reduction reaction is not particularly restricted, but may be an amount of water such that an appropriate amount of the reducing agent, namely hyposulfurous acid or a salt thereof, can be dissolved therein. In general, for example, it is advisable that the amount of the hyposulfurous acid or a salt be adjusted usually to not more than 30 w/w%, and preferably not more than 20 w/w%, relative to the weight of water. From the productivity viewpoint, among others, it is

advisable that the amount be adjusted generally to not less than 1 w/w%, preferably not less than 5 w/w%, and more preferably not less than 10 w/w%.

[0013] The above reduction reaction is carried out in the coexistence of a salt and/or under deoxygenated atmosphere, and at pH of 7 or below. In other words, the above reduction reaction may be carried out under oxygen-containing atmosphere when under in the coexistence of a salt and at pH of 7 or below. Moreover, a condition with no existence of salts is allowable when the reduction is carried out under deoxygenated atmosphere and at pH of 7 or below. Furthermore, it is also possible to carry out the reduction in the existence of salt and under deoxygenated atmosphere, at pH of 7 or below.

[0014] The above salt is not particularly restricted as long as the reduced coenzyme Q_{10} is not oxidized therewith. For example, there may be mentioned a salt constituted from an alkaline metal such as lithium, sodium and potassium, or alkaline earth metals such as magnesium and potassium, with a halogen atom such as fluorine, chloride and bromine, or a residue obtained by excluding a proton from an inorganic acid such as sulfuric acid or an organic acid such as formic acid, acetic acid and propionic acid. Among these, inorganic salts are preferred, and sodium chloride, potassium chloride, sodium sulfate, and the like are more preferred.

[0015] Regarding a concentration of the above salt, high concentration is preferable. Specifically, the concentration is preferably 3 w/w% or above, more preferably 5 w/w% or above, and still more preferably 10 w/w% or above, relative to water. Moreover, it is particularly preferable to dissolve the above salt in a reaction system (aqueous medium) so as to be saturation or close to saturation.

[0016] The above deoxygenated atmosphere can be attained by substitution with an inert gas, pressure reduction, boiling, or a combination of these. It is preferable to carry out at least the substitution with an inert gas, namely to use an inert gas atmosphere. As the inert gas, there may be mentioned, for example, nitrogen gas, carbon dioxide gas, helium gas, argon gas, hydrogen gas and the like. Nitrogen gas is preferred, however.

[0017] It was found that the reduction reaction in the coexistence of a salt and/or under deoxygenated atmosphere mentioned above was particularly effective when hyposulfurous acid or a salt is used as a reducing agent, and that such a reaction greatly contributed to an improvement of the yield in the reduction reaction or decrease of the reducing agent.

[0018] Moreover, the above reduction reaction is carried out at pH of 7 or below, preferably at pH range of 3 to 7, and more preferably at pH range of 3 to 6. The above pH may be adjusted with acids (e.g. mineral acids such as hydrochloric acid and sulfuric acid) or bases (e.g. alkaline metal hydroxides such as sodium hydroxide).

[0019] As described above, various factors may be appropriately controlled for minimizing a residence of oxidized coenzyme Q_{10} or a formation of the oxidized coenzyme Q_{10} as a by-product from the reduced coenzyme Q_{10} , and thus high-quality reduced coenzyme Q_{10} may be synthesized in a high yield.

[0020] In the above reduction reaction, preferable environment are provided, which enable the reduction reaction to preferably proceed, and residence, by-product formation and immixture of the oxidized coenzyme Q_{10} to minimize. Therefore, high yield may be stably attained. Moreover, it is also possible to minimize the amount of the above hyposulfurous acid or salt to be used as a reducing agent.

[0021] The amount of the hyposulfurous acid or salt to be used is not particularly restricted. From the economical viewpoint, however, the amount to be used may be not larger than the charged weight of oxidized coenzyme Q_{10} . The lower limit may be preferably not smaller than about 1/5 by weight, more preferably not smaller than about 2/5 by weight, and still more preferably not smaller than about 3/5 by weight, based on the charged weight of oxidized coenzyme Q_{10} . Thus, the reaction can be more favorably carried out with an amount within the range of about 2/5 by weight of the above-mentioned charged weight to a weight roughly equal to that of the charged weight.

[0022] The above reduction reaction is preferably carried out under forced flowing. The power required for stirring to cause such flowing per unit volume is generally not less than about 0.01 kW/m³, preferably not less than about 0.1 kW/m³, and more preferably not less than about 0.3 kW/m³. The above forced flowing is generally caused by the turning of a stirring blade (s), but the use of a stirring blade (s) is not always necessary if the above flowing can be otherwise obtained. For example, a method based on liquid circulation may be utilized.

[0023] The temperature for the above reduction reaction is not particularly restricted, but preferably 100°C or below, more preferably 80°C or below, and still more preferably 60°C or below. The lower limit of the temperature is preferably the solidification temperature of the system. The reaction may be preferably carried out at a temperature range of 0 to 100°C, more preferably at 0 to 80°C, and still more preferably at 0 to 60°C.

[0024] The substrate concentration in the above reduction reaction is not particularly restricted but the weight of oxidized coenzyme Q_{10} relative to the solvent weight is preferably not less than 1 w/w%, more preferably not less than 3 w/w%, still more preferably not less than 10 w/w%, and particularly preferably not less than 15 w/w%. The upper limit is not particularly restricted, too, but is preferably not more than about 60 w/w%, more preferably not more than 50 w/w%, still more preferably not more than 40 w/w%, and particularly preferably not more than 30 w/w%. Thus, the reaction can be favorably carried out at a substrate concentration of about 1 to 60 w/w%, preferably about 3 to 50 w/w%, and more preferably about 10 to 40 w/w%.

[0025] The above reduction reaction is carried out in an aqueous medium. The aqueous medium may be simple water, or may be a combination of water and an organic solvent.

[0026] The above organic solvent is not particularly restricted, but preferably at least one species selected from hydrocarbons, fatty acid esters, ethers and nitriles in view of the yield and qualities of the reduced coenzyme Q₁₀, and among them, hydrocarbons are preferred. The above organic solvents are effective solvents having great ability to inhibit residence, by-products formation and immixture of the oxidized coenzyme Q₁₀.

[0027] The hydrocarbons are not particularly restricted, but there may be mentioned, for example, aliphatic hydrocarbons, aromatic hydrocarbons, halogenated hydrocarbons, etc. Preferred are aliphatic hydrocarbons and aromatic hydrocarbons, and, among them, aliphatic hydrocarbons are particularly preferred.

[0028] The aliphatic hydrocarbons are not particularly restricted, and may be cyclic or acyclic, or saturated or unsaturated. However, generally they contain 3 to 20 carbon atoms, and preferably 5 to 12 carbon atoms.

[0029] As specific examples, there may be mentioned, for example, propane, butane, isobutane, pentane, 2-methylbutane, cyclopentane, 2-pentene, hexane, 2-methylpentane, 2,2-dimethylbutane, 2,3-dimethylbutane, methylcyclopentane, cyclohexane, 1-hexene, cyclohexene, heptane, 2-methylhexane, 3-methylhexane, 2,3-dimethylpentane, 2,4-dimethylpentane, methylcyclohexane, 1-heptene, octane, 2,2,3-trimethylpentane, isooctane, ethylcyclohexane, 1-octene, nonane, 2,2,5-trimethylhexane, 1-nonene, decane, 1-decene, p-menthane, undecane, dodecane, etc.

[0030] Among them, saturated aliphatic hydrocarbons having 5 to 8 carbon atoms are more preferred, and preferably used are pentane, 2-methylbutane and cyclopentane, which have 5 carbon atoms (referred to as "pentanes"); hexane, 2-methylpentane, 2,2-dimethylbutane, 2,3-dimethylbutane, methylcyclopentane, cyclohexane, which have 6 carbon atoms (referred to as "hexanes"); heptane, 2-methylhexane, 3-methylhexane, 2,3-dimethylpentane, 2,4-dimethylpentane, methylcyclohexane, which have 7 carbon atoms (referred to as "heptanes"); octane, 2,2,3-trimethylpentane, isooctane, ethylcyclohexane, which have 8 carbon atoms (referred to as octanes); and a mixture of these. In particular, the above heptanes are particularly preferred since they have a tendency to show a very high effect to protect reduced coenzyme Q₁₀ from oxidization, and heptane is most preferred.

[0031] The aromatic hydrocarbons are not particularly restricted, but generally they contain 6 to 20 carbon atoms, preferably 6 to 12 carbon atoms, and more preferably 7 to 10 carbon atoms. As specific examples, there may be mentioned, for example, benzene, toluene, xylene, o-xylene, m-xylene, p-xylene, ethylbenzene, cumene, mesitylene, tetralin, butylbenzene, p-cymene, cyclohexylbenzene, diethylbenzene, pentylbenzene, dipentylbenzene, dodecylbenzene, styrene, etc. Preferred are toluene, xylene, o-xylene, m-xylene, p-xylene, ethylbenzene, cumene, mesitylene, tetralin, butylbenzene, p-cymene, cyclohexylbenzene, diethylbenzene and pentylbenzene. More preferred are toluene, xylene, o-xylene, m-xylene, p-xylene, cumene and tetralin, and most preferred is cumene.

[0032] The halogenated hydrocarbons are not particularly restricted, and may be cyclic or acyclic, or saturated or unsaturated. However, acyclic halogenated hydrocarbons are preferably used. More preferred are chlorinated hydrocarbons and fluorinated hydrocarbons, and chlorinated hydrocarbons are still more preferred. Additionally, ones containing 1 to 6 carbon atoms, preferably 1 to 4 carbon atoms, and more preferably 1 to 2 carbon atoms are used.

[0033] As specific examples, for example, there may be mentioned dichloromethane, chloroform, carbon tetrachloride, 1,1-dichloroethane, 1,2-dichloroethane, 1,1,1-trichloroethane, 1,1,2-trichloroethane, 1,1,1,2-tetrachloroethane, 1,1,1,2,2-pentachloroethane, hexachloroethane, 1,1-dichloroethylene, 1,2-dichloroethylene, trichloroethylene, tetrachloroethylene, 1,2-dichloropropane, 1,2,3-trichloropropane, chlorobenzene, 1,1,1,2-tetrafluoroethane, etc.

[0034] Preferred are dichloromethane, chloroform, carbon tetrachloride, 1,1-dichloroethane, 1,2-dichloroethane, 1,1,1-trichloroethane, 1,1,2-trichloroethane, 1,1-dichloroethylene, 1,2-dichloroethylene, trichloroethylene, chlorobenzene and 1,1,1,2-tetrafluoroethane. More preferred are dichloromethane, chloroform, 1,2-dichloroethylene, trichloroethylene, chlorobenzene and 1,1,1,2-tetrafluoroethane.

[0035] The fatty acid esters are not particularly restricted, but there may be mentioned, for example, propionates, acetates, formates, etc. Preferred are acetates and formates, and more preferred are acetates. Ester functional groups thereof are not particularly restricted, but alkyl esters having 1 to 8 carbon atoms, aralkyl esters having 1 to 8 carbon atoms, preferred are alkyl esters having 1 to 6 carbon atoms, and more preferred are alkyl esters having 1 to 4 carbon atoms are used.

[0036] As the propionates, there may be mentioned, for example, methyl propionate, ethyl propionate, butyl propionate, isopentyl propionate, etc.

[0037] As the acetates, there may be mentioned, for example, methyl acetate, ethyl acetate, propyl acetate, isopropyl acetate, butyl acetate, isobutyl acetate, sec-butyl acetate, pentyl acetate, isopentyl acetate, sec-hexyl acetate, cyclohexyl acetate, benzyl acetate, etc. Preferred are methyl acetate, ethyl acetate, propyl acetate, isopropyl acetate, butyl acetate, isobutyl acetate, sec-butyl acetate, pentyl acetate, isopentyl acetate, sec-hexyl acetate and cyclohexyl acetate. More preferred are methyl acetate, ethyl acetate, propyl acetate, isopropyl acetate, butyl acetate and isobutyl acetate. Most preferred is ethyl acetate.

[0038] As the formates, there may be mentioned, for example, methyl formate, ethyl formate, propyl formate, iso-

propyl formate, butyl formate, isobutyl formate, sec-butyl formate, pentyl formate, etc. Preferred are methyl formate, ethyl formate, propyl formate, butyl formate, isobutyl formate and pentyl formate, and most preferred is ethyl formate.

[0039] The ethers are not particularly restricted, and may be cyclic or acyclic, or saturated or unsaturated. But saturated ones are preferably used. Generally, ones containing 3 to 20 carbon atoms, and preferably 4 to 12 carbon atoms and more preferably 4 to 8 carbon atoms are used.

[0040] As specific examples, there may be mentioned, for example, diethyl ether, methyl tert-butyl ether, dipropyl ether, diisopropyl ether, dibutyl ether, dihexyl ether, ethyl vinyl ether, butyl vinyl ether, anisol, phenetole, butyl phenyl ether, methoxytoluene, dioxane, furan, 2-methylfuran, tetrahydrofuran, tetrahydropyran, ethylene glycol dimethyl ether, ethylene glycol diethyl ether, ethylene glycol dibutyl ether, ethylene glycol monomethyl ether, ethylene glycol monoethyl ether, ethylene glycol monobutyl ether, etc.

[0041] Preferred are diethyl ether, methyl tert-butyl ether, dipropyl ether, diisopropyl ether, dibutyl ether, dihexyl ether, anisol, phenetole, butyl phenyl ether, methoxytoluene, dioxane, 2-methylfuran, tetrahydrofuran, tetrahydropyran, ethylene glycol dimethyl ether, ethylene glycol diethyl ether, ethylene glycol dibutyl ether, ethylene glycol monomethyl ether and ethylene glycol monoethyl ether. More preferred are diethyl ether, methyl tert-butyl ether, anisol, dioxane, tetrahydrofuran, ethylene glycol monomethyl ether and ethylene glycol monoethyl ether. More preferred are diethyl ether, methyl tert-butyl ether, anisol, etc., and most preferred is methyl tert-butyl ether.

[0042] The nitriles are not particularly restricted, and may be cyclic or acyclic, or saturated or unsaturated. However, saturated ones are preferably used. Generally, ones containing 2 to 20 carbon atoms, preferably 2 to 12 carbon atoms, and more preferably 2 to 8 carbon atoms are used.

[0043] As specific examples, there may be mentioned, for example, acetonitrile, propionitrile, malonitrile, butyronitrile, isobutyronitrile, succinonitrile, valeronitrile, glutaronitrile, hexanenitrile, heptylcyanide, octylcyanide, undecanenitrile, dodecanenitrile, tridecanenitrile, pentadecanenitrile, stearonitrile, chloroacetonitrile, bromoacetonitrile, chloropropionitrile, bromopropionitrile, methoxyacetonitrile, methyl cyanoacetate, ethyl cyanoacetate, tolunitrile, benzonitrile, chlorobenzonitrile, bromobenzonitrile, cyanobenzoic acid, nitrobenzonitrile, anisonitrile, phthalonitrile, bromotolunitrile, methyl cyanobenzoate, methoxybenzonitrile, acetylbenzonitrile, naphthonitrile, biphenylcarbonitrile, phenylpropionitrile, phenylbutyronitrile, methylphenylacetoneitrile, diphenylacetoneitrile, naphthylacetoneitrile, nitrophenylacetoneitrile, chlorobenzylcyanide, cyclopropanecarbonitrile, cyclohexanecarbonitrile, cycloheptanecarbonitrile, phenylcyclohexanecarbonitrile, tolylcyclohexanecarbonitrile, etc.

[0044] Preferred are acetonitrile, propionitrile, succinonitrile, butyronitrile, isobutyronitrile, valeronitrile, methyl cyanoacetate, ethyl cyanoacetate, benzonitrile, tolunitrile and chloropropionitrile. More preferred are acetonitrile, propionitrile, butyronitrile and isobutyronitrile, and most preferred is acetonitrile.

[0045] Among the above organic solvents, it is preferred to use a solvent with low miscibility to water. This makes it possible to adequately carry out the above reduction reaction and, additionally, post-treatments after the reduction reaction.

[0046] In selecting the solvent to be used from among the solvents mentioned above, such properties as boiling point and viscosity are preferably taken into consideration; for example, the solvent should have a boiling point which allows appropriate warming for increasing the solubility and facilitates solvent recovery from crystallization filtrates and a solvent removal from wet masses by drying (about 30 to 150°C at 1 atm), a melting point such that solidification hardly occurs in handling at room temperature as well as upon cooling to room temperature or below (not lower than about 0°C, preferably not lower than about 10°C, still more preferably not lower than about 20°C), and a low viscosity (not higher than about 10 cp at 20°C). From the industrial operation viewpoint, a solvent which is hardly volatile at ordinary temperature is preferred; generally, for example, one having a boiling point of not lower than about 80°C is preferred, and one having a boiling point of not lower than about 90°C is particularly preferred.

[0047] The above reduction reaction can be driven to completion usually within 5 hours, preferably within 3 hours, and more preferably within 1 hour.

[0048] The generated reduced coenzyme Q₁₀ is extracted into an organic solvent from thus-obtained aqueous mixture after the reduction reaction to recover an organic phase containing the reduced coenzyme Q₁₀. Then, if necessary (preferably), said organic phase is further washed with water repeatedly to completely remove impurities. The water to be used for washing is not particularly restricted, but preferably water or an aqueous solution containing a salt, preferably an inorganic salt such as sodium chloride, potassium chloride and sodium sulfate, etc. in view of easiness of liquid separation (wherein, concentration of the salt is preferably high, and it is usually 5 w/w% or above, preferably about 10 w/w% or above, and more preferably a concentration of saturation or close to saturation). The extraction and washing mentioned above may be carried out under acidic condition, preferably at pH of 6 or below, and more preferably at pH of 5 or below for minimizing the formation of oxidized coenzyme Q₁₀ as a by-products.

[0049] The organic solvent to be used for the extraction mentioned above is not particularly restricted. But from the fore mentioned viewpoints, it is preferable to use one species selected from hydrocarbons, fatty acid esters, ethers and nitriles as mentioned above. When the organic solvent is used together in the above reduction reaction, the same organic solvent is preferably used also as an extraction solvent.

[0050] The thus-obtained organic phase containing the reduced coenzyme Q₁₀ may be then subjected to operations appropriately combined among cooling, concentration, solvent substitution or the like, thereby crystallizing reduced coenzyme Q₁₀. The high-quality reduced coenzyme Q₁₀ recovered by the above method may be dried under normal pressure or reduced pressure.

[0051] The above-mentioned treatments after the reduction reaction, namely a series of operation from extraction to recovering dried crystal, are carried out under deoxygenated atmosphere. Preferably, the treatment may be carried out, for example, under inert gas atmosphere such as nitrogen gas, helium gas, carbon dioxide gas, argon gas and hydrogen gas atmosphere, and particularly preferably under nitrogen gas atmosphere.

[0052] In accordance with the present invention, various factors to inhibit residence, by-product formation and im-mixture of the oxidized coenzyme Q₁₀ can be appropriately controlled, and thus high-quality reduced coenzyme Q₁₀ may be obtained in a convenient and efficient manner at a high yield. The reduced coenzyme Q₁₀ as obtained in accordance with the present invention is a product with exceedingly high-quality, and can be expected to have a reduced coenzyme Q₁₀/oxidized coenzyme Q₁₀ weight ratio of not lower than 96/4, preferably not lower than 98/2, more preferably not lower than 99/1.

BEST MODE FOR CARRYING OUT THE INVENTION

[0053] The following examples illustrate the present invention in further detail. These examples are, however, by no means limitative of the scope of the present invention. In the examples, the purity of reduced coenzyme Q₁₀ and the reduced coenzyme Q₁₀/oxidized coenzyme Q₁₀ weight ratio were determined by the HPLC analysis specified below. The reduced coenzyme Q₁₀ purity values as determined, however, are by no means indicative of the limit purity value attainable in accordance with the present invention. Likewise, the reduced coenzyme Q₁₀/oxidized coenzyme Q₁₀ weight ratio values obtained never indicate the upper limit to that ratio.

(HPLC conditions)

[0054] Column: SYMMETRY C18 (product of Waters), 250 mm (in length), 4.6 mm (in inside diameter); mobile phase: C₂H₅OH:CH₃OH = 4:3 (v/v); detection wavelength: 210 nm; flow rate: 1 ml/min; retention time of reduced coenzyme Q₁₀: 9.1 min; retention time of oxidized coenzyme Q₁₀: 13.3 min.

(Example 1)

[0055] While stirring (stirring power consumption 0.3 kW/m³), 100 g of the oxidized coenzyme Q₁₀ (purity 99.4%) at 48°C, a aqueous solution prepared by dissolving 80 g of sodium hyposulfite (purity 75% or more) in 1100 g of 10 w/w% brine was gradually added as a reducing agent, to carry out a reduction reaction at 48°C and pH of 4 to 6. After the lapse of 2 hours, 1000 g of heptane was added thereto and an aqueous phase was removed. Then, a heptane phase was washed for 6 times with 1000 g of saturated brine adjusted to pH of 3 by hydrochloric acid, to give a heptane solution of the reduced coenzyme Q₁₀. All the above operations were carried out under nitrogen atmosphere. The weight ratio of the reduced coenzyme Q₁₀/oxidized coenzyme Q₁₀ in the heptane solution was 99.5/0.5, and the yield of the reduced coenzyme Q₁₀ was 99% by mole.

(Example 2)

[0056] A heptane solution of reduced coenzyme Q₁₀ was obtained by the same procedure as in Example 1 except that reduction reaction was carried out in the atmosphere. The reduced coenzyme Q₁₀/oxidized coenzyme Q₁₀ weight ratio in the heptane solution was 99.0/1.0, and yield of the reduced coenzyme Q₁₀ was 99% by mole.

(Example 3)

[0057] A heptane solution of reduced coenzyme Q₁₀ was obtained by the same procedure as in Example 1 except that an aqueous solution (no sodium chloride added) prepared by dissolving 80 g of sodium hyposulfite (purity 75% or more) in 1000 g of water was used as a reducing agent. The weight ratio of reduced coenzyme Q₁₀/oxidized coenzyme Q₁₀ in the heptane solution was 99.4/0.6, and the yield of the reduced coenzyme Q₁₀ was 99% by mole.

(Comparative Example 1)

[0058] A heptane solution of the reduced coenzyme Q₁₀ was obtained by the same procedure as in Example 1 except that an aqueous solution (no sodium chloride added) prepared by dissolving 80 g of sodium hyposulfite (purity 75% or

more) in 1000 g of water was used as a reducing agent and the reduction reaction was carried out in the atmosphere. The weight ratio of reduced coenzyme Q₁₀/oxidized coenzyme Q₁₀ in the heptane solution was 87.4/12.6, and the yield of the reduced coenzyme Q₁₀ was 87% by mole.

(Example 4)

[0059] Oxidized coenzyme Q₁₀ (100 g; purity 99.4%) was dissolved in 1000 g of heptane at 25°C. While stirring (stirring power consumption 0.3 kW/m³), an aqueous solution prepared by dissolving 62 g of sodium hyposulfite (purity 75% or more) in 1100 g of 10 w/w% brine was gradually added as a reducing agent, to carry out the reduction reaction at 25°C and pH of 4 to 6. After the lapse of 2 hours, an aqueous phase was removed from the reaction solution, and a heptane phase was washed for 6 times with 1000 g of saturated brine adjusted to pH of 3 by hydrochloric acid, to give a heptane solution of reduced coenzyme Q₁₀. All the above operations were carried out under nitrogen atmosphere. The weight ratio of reduced coenzyme Q₁₀/oxidized coenzyme Q₁₀ in the heptane solution was 99.5/0.5, and the yield of the reduced coenzyme Q₁₀ was 99% by mole.

(Example 5)

[0060] A heptane solution of the reduced coenzyme Q₁₀ was obtained by the same procedure as in Example 4 except that the reduction reaction was carried out in the atmosphere. The weight ratio of the reduced coenzyme Q₁₀/oxidized coenzyme Q₁₀ in the heptane solution was 99.3/0.7, and the yield of the reduced coenzyme Q₁₀ was 99% by mole.

(Example 6)

[0061] A heptane solution of the reduced coenzyme Q₁₀ was obtained by the same procedure as in Example 4 except that an aqueous solution (no sodium chloride added) prepared by dissolving 62 g of sodium hyposulfite (purity 75% or more) in 1000 g of water was used as a reducing agent. The weight ratio of the reduced coenzyme Q₁₀/oxidized coenzyme Q₁₀ in the heptane solution was 99.4/0.6, and the yield of the reduced coenzyme Q₁₀ was 99% by mole.

(Example 7)

[0062] A hexane solution of reduced coenzyme Q₁₀ was obtained by the same procedure as in Example 5 except that hexane was used as a solvent for dissolving the oxidized coenzyme Q₁₀. The weight ratio of reduced coenzyme Q₁₀/oxidized coenzyme Q₁₀ in the hexane solution was 99.1/0.9, and the yield of reduced coenzyme Q₁₀ was 99% by mole.

(Comparative Example 2)

[0063] A heptane solution of the reduced coenzyme Q₁₀ was obtained by the same procedure as in Example 4 except that an aqueous solution (no sodium chloride added) prepared by dissolving 62 g of sodium hyposulfite (purity 75% or more) in 1000 g of water was used as a reducing agent and the reduction reaction was carried out in the atmosphere. The weight ratio of reduced coenzyme Q₁₀/oxidized coenzyme Q₁₀ in the heptane solution was 91.0/9.0, and the yield of the reduced coenzyme Q₁₀ was 91% by mole.

(Example 8)

[0064] A hexane solution of reduced coenzyme Q₁₀ was obtained by the same procedure as in Example 4 except that hexane was used as a solvent for dissolving oxidized coenzyme Q₁₀ and a solution (no sodium chloride added) dissolving 60 g of sodium hyposulfite (purity 75% or more) in 1000 g of water was used as a reducing agent. The weight ratio of reduced coenzyme Q₁₀/oxidized coenzyme Q₁₀ in the hexane solution was 99.3/0.7, and the yield of the reduced coenzyme Q₁₀ was 99% by mole.

(Comparative Example 3)

[0065] A heptane solution of reduced coenzyme Q₁₀ was obtained by the same procedure as in Example 8 except that reduction reaction was carried out in the atmosphere. The weight ratio of reduced coenzyme Q₁₀/oxidized coenzyme Q₁₀ in the heptane solution was 90.9/9.1, and yield of the reduced coenzyme Q₁₀ was 91% by mole.

(Example 9)

[0066] A heptane solution of the reduced coenzyme Q_{10} was obtained by the same procedure as in Example 1 except that an aqueous solution prepared by dissolving 80 g of sodium hyposulfite (purity 75% or more) in 1050 g of 5 w/w% brine was used as a reducing agent and the reduction reaction was carried out in the atmosphere. The weight ratio of reduced coenzyme Q_{10} /oxidized coenzyme Q_{10} was 98.9/1.1, and the yield of the reduced coenzyme Q_{10} was 99% by mole.

(Comparative Example 4)

[0067] A heptane solution of reduced coenzyme Q_{10} was obtained by the same procedure as in Example 4 except that the reduction reaction was carried out at pH range of 8 to 9. The weight ratio of reduced coenzyme Q_{10} / oxidized coenzyme Q_{10} in the heptane solution was 54.0/46.0, and the yield of the reduced coenzyme Q_{10} was 54% by mole.

(Reference Example 1)

[0068] One gram of reduced coenzyme Q_{10} (weight ratio of reduced coenzyme Q_{10} /oxidized coenzyme Q_{10} is 99.6/0.4) was dissolved in 20 g of various solvents shown in Table 1 at 25°C. In the atmosphere, the weight ratio of reduced coenzyme Q_{10} /oxidized coenzyme Q_{10} in the solutions were measured after stirring for 24 hours at 35°C. The results are shown in Table 1.

Table 1

Solvent	R
Heptane	99.1/0.9
Hexane	98.7/1.3
Toluene	98.8/1.2
Chloroform	98.9/1.1
Ethyl acetate	98.9/1.1
Methyl tert-butyl ether	98.6/1.4
Tetrahydrofuran	98.5/1.5
R: Reduced coenzyme Q_{10} / Oxidized coenzyme Q_{10} weight ratio	

(Reference Example 2)

[0069] One gram of reduced coenzyme Q_{10} (weight ratio of reduced coenzyme Q_{10} /oxidized coenzyme Q_{10} is 99.6/0.4) was dissolved in 100 g of various solvents shown in Table 2 at 35°C. In the atmosphere, the weight ratio of reduced coenzyme Q_{10} /oxidized coenzyme Q_{10} in the solutions were measured after stirring for 24 hours at 25°C. The results are shown in Table 2.

Table 2

Solvent	R
Heptane	96.7/3.3
Ethyl acetate	96.4/3.6
Acetonitrile	96.0/4.0
R: Reduced coenzyme Q_{10} / Oxidized coenzyme Q_{10} weight ratio	

INDUSTRIAL APPLICABILITY

[0070] Since the present invention has the constitution mentioned above, high-quality reduced coenzyme Q_{10} may be obtained in a convenient and efficient manner. Thus, the method is suited for the industrial scale production.

Claims

1. A method of producing a reduced coenzyme Q₁₀
which comprises reducing an oxidized coenzyme Q₁₀ in an aqueous medium with the use of hyposulfurous
5 acid or a salt thereof,
said reduction being carried out in the coexistence of a salt and/or under deoxygenated atmosphere, and at
pH of 7 or below.
2. The method according to Claim 1,
10 wherein the salt is an inorganic salt.
3. The method according to Claim 1 or 2,
wherein a concentration of the salt in water is 3 w/w% or more.
4. The method according to any one of Claims 1 to 3,
15 wherein the deoxygenated atmosphere is an inert gas atmosphere.
5. The method according to any one of Claims 1 to 4,
wherein pH range is between 3 and 7.
6. The method according to any one of Claims 1 to 5,
20 wherein the aqueous medium is a mixed medium comprising water and an organic solvent.
7. The method according to Claim 6,
25 wherein the organic solvent is at least one species selected from the group consisting of hydrocarbons, fatty
acid esters, ethers and nitriles.
8. The method according to Claim 6,
wherein the organic solvent is a hydrocarbon.
9. The method according to Claim 6,
30 wherein the organic solvent is one of pentanes, hexanes, heptanes or octanes.
10. The method according to Claim 6,
35 wherein the organic solvent is one of heptanes.
11. The method according to any one of Claims 1 to 10,
wherein an organic phase containing the reduced coenzyme Q₁₀ is recovered by extracting the generated
40 reduced coenzyme Q₁₀ into the organic solvent after the reduction reaction.
12. The method according to Claim 11,
wherein the extraction is carried out at pH of 6 or below.
13. The method according to Claim 11 or 12,
45 wherein the organic phase recovered by the extraction is further washed with water.
14. The method according to any one of Claims 1 to 13 which is carried out under deoxygenated atmosphere.

INTERNATIONAL SEARCH REPORT

International application No.

PCT/JP02/07147

A. CLASSIFICATION OF SUBJECT MATTER
Int.Cl⁷ C07C43/23, 41/26

According to International Patent Classification (IPC) or to both national classification and IPC

B. FIELDS SEARCHED

Minimum documentation searched (classification system followed by classification symbols)
Int.Cl⁷ C07C43/23, 41/26

Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched

Electronic data base consulted during the international search (name of data base and, where practicable, search terms used)
CA (STN), REGISTRY (STN), CASREACT (STN)

C. DOCUMENTS CONSIDERED TO BE RELEVANT

Category*	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
A	JP 53-133687 A (Nisshin Flour Milling Co., Ltd.), 21 November, 1978 (21.11.78), Page 2, upper right column to page 3, upper right column (Family: none)	1-14
A	JP 57-70834 A (Takara Shuzo Kabushiki Kaisha), 01 May, 1982 (01.05.82), Example 1 (Family: none)	1-14
A	WO 98/7417 A1 (Kaneka Corp.), 26 February, 1998 (26.02.98), & US 6184255 B1 & EP 956854 A1 & JP 10-109933 A	1-14

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Date of the actual completion of the international search
21 October, 2002 (21.10.02)

Date of mailing of the international search report
05 November, 2002 (05.11.02)

Name and mailing address of the ISA/
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